Multiple CO₂ capture in stable metal-doped graphene: a theoretical trend study

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Identifying stable systems with high CO₂ adsorption capacity is an essential goal in CO₂ capture and storage technologies. We have carried out a comprehensive first-principles study to explore the CO₂ capture capacity of 16 representative metal-doped graphene systems where the metal dopants can be stabilized by single- and double-vacancies. The maximum number of adsorbed CO₂ molecules was determined by a combination of adsorption energy and bond distance criteria. Generally, while the double-vacancy can bind metal dopants more strongly than the single-vacancy, single-vacancy graphene with metal dopants are better sorbents, with each Ca, Sc and Y dopant binding up to 5 CO₂ molecules. CO₂ capture involves significant charge transfer between the CO₂ molecule and the dopant–vacancy complexes, where defective graphene acts as a charge reservoir for binding CO₂ molecules. Some systems are predicted to involve the formation of a bent CO₂ anion. Ca-doped single- and double-vacancy graphene systems, however, readily form oxides upon reaction with CO₂, thus they are less reusable for CO₂ capture.

1 Introduction

Carbon dioxide capture and storage (CCS) is a group of technologies dedicated to the capture of CO₂ followed by compression, transport, and storage. It is an intense area of research because it represents a key approach for curbing CO₂ emission and reducing the impact of energy production on climate change. CCS implemented in modern power plants is estimated to reduce CO₂ emission by 80–90%. However, currently the capture alone will increase the energy requirements of a plant by 25–40% because existing methods of capture are energy intensive and are not cost effective. Intensive research efforts are directed towards achieving effective and energy-efficient capture technologies.

New alternative materials are being investigated, such as porous coal, activated carbons (AC), metal–organic frameworks (MOFs), zeolites and fullerenes. Recently, low-dimensional carbon-based materials including graphene, graphene oxide and carbon nanotubes (CNTs) have been considered as promising adsorbents for CO₂. As such, they have been widely investigated by various research groups at low and very high CO₂ pressures. This is because of their large surface area, easy-to-design pore structures, surface functionalization and low energy requirements for regeneration. In particular, a graphene sheet with adsorbed gases could be regenerated to its initial state within 100–200 s by annealing at 150 °C in vacuum or short ultra-violet irradiation. Furthermore, it was experimentally demonstrated that using graphene for CO₂ capture, at a pressure of 11 bar and room temperature, has a higher capture capacity than other carbon nanostructures and zeolites. Recently, the reactivity of graphene was also shown to be tunable by varying the external electric field.

Functionalization and doping are known to alter the chemical properties of graphene sheets to enhance their response to gases. Studies show that nitrogen doping of graphene enhances its CO₂ capture capacity. Nanostructured polyaniline decorated graphene was investigated by Mishra et al. and found to have high selectivity, and nitrogen doping was shown to enhance CO₂ capture in a microporous carbon system. Doping with metal adatoms is expected to be even more appealing due to the inherent diversity of their chemical properties. Metal adatoms adsorbed on pristine, single vacancy (SV) and double vacancy (DV) graphene sheets have been studied, on the basis of first principles density functional theory (DFT) calculations, owing to the feasibility of performing fine deposition of adatoms into the graphene surface. Fabrication of DV rich graphene has been achieved experimentally through high energy ion bombardment. Deposition of Pt on graphene was the subject of a recent experimental investigation, and Al was theoretically and experimentally shown to improve the adsorption ability and sensitivity of carbon nanotubes and graphene towards gases. In the present study, we focus on the so-called 585-type DV configuration which is less stable than the reconstructed 555777 configuration by about 0.9 eV. However, to our knowledge, the 555777-type DV has not yet been experimentally fabricated, while the 585 configuration has been produced by high energy ion bombardment. It is possible that...
combining metal species with DV and SV graphene will enhance
their sensitivity and selectivity by favouring adsorption/desorption processes. Theoretical studies show that DV and SV graphene can stabilise metal adatoms and overcome the clustering difficulty, thus facilitating device applications.\textsuperscript{18,25}

It is easy to envisage that different metal dopants may have different binding behavior with CO\textsubscript{2}.\textsuperscript{26} In this regard, the maximum CO\textsubscript{2} adsorption capacity is a key parameter in identifying potential CCS materials. However, to our knowledge, existing \textit{ab initio} studies in the literature consider only one CO\textsubscript{2} molecule capture. Moreover, a systematic trend study is appealing in evaluating different candidate systems for CO\textsubscript{2} capture, yet such an investigation is still lacking. Here, we perform a comprehensive \textit{DFT} trend study of the adsorption capacities of metal adatoms on \textit{SV} and \textit{DV} graphene. We consider multiple CO\textsubscript{2} capture on 16 different metal–graphene systems (metal M = Al, Ca, Co, Cr, Cu, Fe, Mg, Mn, Nb, Ni, Pt, Sc, Ti, V, Y, and Zr). These metallic elements were selected to test the performance of species known to exhibit high sensitivity to gases as well as to explore the capacity of elements that have seldom been explored for CO\textsubscript{2} capture. Moreover, a systematic trend study is appealing in evaluating different candidate systems for CO\textsubscript{2} capture, yet such an investigation is still lacking. Here, we perform a comprehensive \textit{DFT} trend study of the adsorption capacities of metal adatoms on \textit{SV} and \textit{DV} graphene. We consider multiple CO\textsubscript{2} capture on 16 different metal–graphene systems (metal M = Al, Ca, Co, Cr, Cu, Fe, Mg, Mn, Nb, Ni, Pt, Sc, Ti, V, Y, and Zr). These metallic elements were selected to test the performance of species known to exhibit high sensitivity to gases as well as to explore the capacity of elements that have seldom been explored for CO\textsubscript{2} capture to date. In particular, we consider elements of the first row of transition metal (TM) elements which have good sensing properties to CO\textsubscript{2} in their bulk oxide form.\textsuperscript{27} We denote the metal–graphene system with the notation M/DV (metal atom adsorbed on double-vacancy graphene) and M/SV (metal atom adsorbed on single-vacancy graphene).

We focus primarily on the atomic structure, energetics, charge transfer and charge distribution in these systems to determine their CO\textsubscript{2} capture and storage capacity. A few elements among the 16 investigated (namely, Ca, Nb, Sc, Zr and Y) were found to have exceptionally high CO\textsubscript{2} adsorption capability. Our study reveals that CO\textsubscript{2} capture involves significant charge transfer between the CO\textsubscript{2} molecule and the dopant–vacancy complexes, where defective graphene acts as a charge reservoir for binding CO\textsubscript{2} molecules. Interestingly, the CO\textsubscript{2} anion forms in the Ca and Nb systems where the first adsorbed CO\textsubscript{2} molecule experiences significant bending, and acquires a negative charge.

### 2 Calculation details

We perform spin-unrestricted density functional theory (DFT) calculations with the SIESTA code,\textsuperscript{28} using the generalized gradient approximation (GGA) for the exchange–correlation functional as developed by Perdew, Burke and Ernzerhof (PBE).\textsuperscript{29} SIESTA uses basis sets comprised of numerical atomic orbitals, and approximates the atomic potential in terms of Troullier–Martins\textsuperscript{30} norm-conserving pseudopotentials. The auxiliary basis set uses a real-space mesh with a kinetic energy cutoff of 500 Ry, and the basis functions are radially confined using an energy shift of 0.005 Ry (see ref. 28 for details). Our simulation consists of a rectangular 72-atom graphene sheet with lattice parameters $a = 12.896$ Å and $b = 14.942$ Å. The calculated lattice constant of graphene is 2.48 Å, which is in very good agreement with the experimental value of 2.46 Å (ref. 31) and previous DFT GGA calculations (e.g. 2.44 Å and 2.47 Å from ref. 31 and 32, respectively). A reciprocal-space $k$-point mesh of $5 \times 5 \times 1$ is used. We allow full atomic relaxation until the forces on the atoms are less than 0.03 eV Å\textsuperscript{-1}.

\section{3 Results and discussion}

The adsorption energies of the metal adatom on the defective graphene surface, and the average adsorption energies of CO\textsubscript{2} molecules on metal-doped DV/SV graphene are given below:

$$E_{\text{ad}}^M = E_{\text{M/SV(DV)Gr}} - E_{\text{SV(DV)Gr}} - E_{\text{free}}$$  \hspace{1cm} (1)

$$E_{\text{ad}}^{\text{CO2}} = \left( E_{\text{M/SV(DV)Gr}}^{\text{CO2}} - E_{\text{M/SV(DV)Gr}} - nE_{\text{free}}^{\text{CO2}} \right) / n$$  \hspace{1cm} (2)

where $E_{\text{ad}}^M$ is the adsorption energy of the metal atom on SV/DV graphene, $E_{\text{M/SV(DV)Gr}}$ is the total energy of the M/DV(SV) system, $E_{\text{SV(DV)Gr}}$ is the total energy of the SV(DV) graphene sheet unit cell, $E_{\text{free}}$ is the energy of the free metal atom, $E_{\text{ad}}^{\text{CO2}}$ is the average adsorption energy of $n$ CO\textsubscript{2} molecules on the M/SV(DV) system, $E_{\text{M/SV(DV)Gr}}^{\text{CO2}}$ is the total energy of the M/SV(DV) system with $n$ adsorbed CO\textsubscript{2} molecules and $E_{\text{free}}^{\text{CO2}}$ is the total energy of the free CO\textsubscript{2} molecule (note, we do not include the vacancy formation energy, as typically the defective graphene systems are produced by ion-bombardment).

\subsection{3.1 Metal atom adsorption on SV and DG graphene}

We first investigate the atomic and electronic structure, and energetics of the metal–graphene systems for adsorption on SVs and DVs. The vertical distances of the metal adatoms to the graphene plane, the adsorption energies (eqn (1)) and associated total unit-cell magnetic moments are presented in Fig. 1. The results are in reasonable agreement with values reported in an earlier theoretical work.\textsuperscript{28} Generally, adsorption on DV graphene results in shorter vertical distances because one DV has a larger hole with four edge carbon atoms, while one SV has three edge carbon atoms. Ca atoms adsorb with the highest vertical distance (2.40 Å in Ca/SV, 1.98 Å in Ca/DV) and with the lowest adsorption energy (−1.35 eV in Ca/SV, −0.74 eV in Ca/DV). The 3d TMs bind at vertical distances in the range 1.3 Å to 1.7 Å (except for Sc which is 1.93 Å), while 4d TMs adsorb at relatively larger vertical distances because of the larger ionic radius. The systems Mg/SV and Mg/DV also have relatively low adsorption energy (−1.52 eV in Mg/SV, −1.72 eV for Mg/DV). All of the TMs considered bind relatively strongly with the C atoms with adsorption energies ranging from −7.98 eV for Zr/SV to −3.51 eV for Cu/SV. Comparing with the adatom adsorption energies reported in ref. 33, several metal species are predicted to bind more strongly to DV or SV graphene than to covalent organic frameworks, namely Ti, Sc, Ca and Mg.

The magnetization of the adatom/graphene systems is shown in units of $\mu_B$ on the red background in Fig. 1. The undoped DV/graphene is non-magnetic, and the undoped SV/graphene has a spin moment of 1.62 $\mu_B$ (in agreement with ref. 34). The Ti, Sc, Ni Al, Y, Zr, and Pt adatoms totally quench the spin polarization of the SV/graphene system, while Mn, Ca and Mg increases the spin polarization of the SV/graphene structure the most (2.95 $\mu_B$, 1.95 $\mu_B$ and 1.95 $\mu_B$, respectively).
For the DV/graphene system, doping with Ti, Mg, Zr, Ni and Cu does not magnetize the system, whereas doping with the rest of the considered elements does. Maximum magnetization in doped DV/graphene systems occur in Mn/DV and Fe/DV (3.42 $\text{mB}$ and 3.76 $\text{mB}$, respectively).

We study charge transfer by using the Voronoi deformation density method as implemented in SIESTA. In the metal-doped SV and DV graphene structures considered, the metal atoms always donate charge to the graphene sheet, mainly to the edge C atoms of the vacancies. For each species, the amount of donated charge is similar in SV and DV systems. The largest charge transfer occurs in Pt/SV and Pt/DV systems, where the Pt atom loses 0.78 $\text{e}$ in both systems. The Ni atom for the Ni/DV system loses the least amount of charge (0.073 $\text{e}$), while in the Ni/SV, Co/SV, Fe/SV, the adatom loses electron charge from 0.1e to 0.3e. For the rest of the systems the adatom loses electron charge ranging from 0.3e to 0.6e.

### 3.2 CO$_2$ adsorption on adatom/SV and DV graphene

To ascribe the number of CO$_2$ molecules that a system can bind, we adopt a combination of two criteria: (1) the average adsorption energy (eqn (2)) should be above 0.3 eV, and (2) the distance between the additional adsorbed CO$_2$ molecule and the metal adatom be less than 4 Å (the distance for this purpose is defined as the distance between the metal adatom and the closest atom of the CO$_2$ molecule). We have chosen the 0.3 eV threshold arbitrarily based on the physisorption threshold in the literature, and the value of 4 Å based on the correlation between the average adsorption energies as they approach 0.3 eV and the associated bond distances. We add the $n^\text{th}$ CO$_2$ molecules to the relaxed structure with $(n-1)$ CO$_2$ molecules as long as both criteria are satisfied. In this manner, the adsorption capacities are assigned as given in Fig. 1. In order to confirm that our atomic relaxation calculations have yielded structural minima, we have performed relaxation with different initial configurations for the case of adsorption of two CO$_2$ molecules at Ca/SV, Mg/SV, and Sc/SV. For each species, we obtained identical total energies and consistent converged structures giving confidence in our approach.

Among the systems considered, the distance between the metal adatom and the bonded CO$_2$, the shortest distance is for Co/SV (2.02 Å), Fe/SV (2.00 Å) and Ni/SV (2.05 Å). For the systems with highest capture capacities (Ca, Nb, Y, Sc, Y and Zr), the distances are within the range 2.2–2.7 Å. An interesting aspect of CO$_2$ adsorption on M/SV and M/DV systems is the angle of inclination of the CO$_2$ molecule to the graphene surface. For some systems with collinear CO$_2$, the CO$_2$ molecule aligns almost parallel to the plane of the graphene sheet. These systems are Fe/DV, Cu/DV, Mn/DV, Ni/DV, Pt/DV, Ti/DV, V/DV and Y/DV. Interestingly, Mg/DV, Mg/SV and Pt/SV are predicted to align the CO$_2$ molecule vertically (Fig. 2).

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**Fig. 1** For each element, the white box shows the maximum adsorption capacities of SV and DV systems. The blue box shows the vertical distance (relative to the plane of graphene, in Å) of the M adatom and the red box shows the total magnetic spin moment (in $\text{mB}$). The bottom grey row of the box shows the adsorption energy of the metal adatom on SV/DV graphene. The vertical distance and total magnetic moment are for the doped systems before adsorbing the first CO$_2$ (at left side of the arrow, the “pure” system) and after adsorption of the maximum number of CO$_2$ molecules (at the right side of the arrow, the “full” system).
Based on the calculated average adsorption energy, the following systems bond very weakly to the first adsorbed CO2 molecule, and hence, do not have good potential as CO2 sorbent systems: Al/DV (−0.20 eV), Ti/DV (−0.08 eV), Cu/DV (−0.12 eV), Pt/DV (−0.09 eV), V/DV (−0.08 eV), Mn/DV (−0.08 eV), Fe/DV (−0.07 eV), Y/DV (−0.09 eV), Co/DV (−0.02 eV), Ni/DV (−0.13 eV), V/SV (−0.04 eV), Pt/SV (−0.23 eV) and Cu/SV (−0.26 eV). Thus, for these systems, we set the CO2 capacity to be zero (Fig. 1). Fig. 3 shows the average adsorption energy $E_{ad}^{CO2}$ as a function of the number of CO2 molecules adsorbed in systems which can bind more than one CO2 molecule. Results are shown for adsorption on both the SV (Fig. 3(a)) and DV (Fig. 3(b)) structures. The maximum number of adsorbed CO2 molecules are as follows: Ca/DV (5), Ca/SV (5), Nb/DV (4), Nb/SV (4), Sc/DV (4), Sc/SV (5), Cr/SV (4), Zr/DV (2), Zr/SV (4), Y/SV (5), Ti/SV (4), Mn/SV (3) and Fe/SV (2). For most systems studied, the average adsorption energy decreases as $n$ increases, except for Y/SV and Mn/SV. For these two systems, the adsorption is rather weak, with the average adsorption energy around 0.3 eV, resulting in a rather flat adsorption energy variation with increasing CO2.

Ca is a non-TM adatom with a remarkably strong CO2 capture performance, and the capacity to adsorb up to 5 (Ca/SV) and 5 (Ca/DV) CO2 molecules. Ca-doped CNTs and pristine graphene was investigated in ref. 37. For the purpose of comparison with the results of ref. 37, we consider the adsorption of Ca on pristine graphene in a 72 C atom supercell, and its interaction with a single CO2 molecule. The binding energy of Ca is −0.81 eV. This value agrees reasonably well with the result of ref. 37 of −1.02 eV (we note that the difference in value may be due to the different sized supercells used: ours contained 72 C atoms, while theirs contained 32 C atoms). However, Ca-doped graphene might not be the best system for sustainable CO2 capture because of the strong tendency of Ca to leave its vacancy site and bond with an O atom to form CaO clusters, accompanied by the formation of CO gas. This is demonstrated using the following equation,17

$$\Delta E^{MO} = E(MO) + E(CO) + E_{M/SV(DV)/Gr} - E_{MO2}^{CO2}$$ (3)

$$\Delta E^{MO_2} = E(MO_2) + 2E(CO) + E_{M/SV(DV)/Gr} - E_{MO2}^{CO2}$$ (4)

Here $\Delta E^{MO}$ is the formation energy if the metal oxide cluster, $\Delta E^{MO_2}$ is the formation energy of the metal dioxide cluster, $E(MO)/E(MO_2)$ is the total energy of the bulk oxide and $E(CO)$ is the total energy of the CO molecule. The obtained values are $\Delta E^{MO} = −0.9$ eV for Ca/SV, and $\Delta E = −1.33$ eV for Ca/DV. For comparison, $\Delta E^{MO} = +2.45$ eV for Nb/SV, $\Delta E^{MO} = +1.02$ eV for Ti/SV (forming TiO), $\Delta E = +1.29$ eV Ti/SV (forming TiO) and $\Delta E^{MO} = +0.89$ eV for Zr/SV. These positive numbers indicate that TM atoms are more suitable for CCS applications than Ca because the doped graphene system is stable against clustering into its oxide upon CO2 adsorption. It is worth noting that although the formation energies for Ca/SV and Ca/DV have negative signs, those two structures might still have a potential advantage because CaO + CO2 $\rightarrow$ CaCO3 is a thermodynamically favorable process.18 Therefore, out of the structures with multiple CO2 capacity, the systems Nb/SV, Nb/DV, Sc/CV, Zr/SV, Y/SV, Ti/SV, Cr/SV and Mn/SV appear to be the most promising candidates for CCS applications as they can bind multiple CO2 molecules. Interestingly, it can be seen that the elements considered in this study display a trend in their capture capacity (cf. Fig. 1): the capacity roughly increases as we go to the left of the periodic table (except for vanadium) for both first and second row TM elements.

To gain insight into the nature of the bond between CO2 and the TM/graphene system, the electronic properties of the systems are investigated by studying the density of states (DOS) and partial DOS (pDOS). As an example, we consider Nb/SV, where the results are shown in Fig. 4. The corresponding atomic structure is shown in Fig. 4(e). In Fig. 4(a), the total DOS
are shown for the Nb/SV structure without CO$_2$, and in Fig. 4(b) the total DOS with the adsorption of one CO$_2$ molecule. While pristine graphene is a semimetal, Nb/SV graphene is clearly metallic. When CO$_2$ bonds to the Nb atom, the O-related 2p states of CO$_2$ hybridize with the 4d states of the Nb atom, which is the reason for the relatively high adsorption energy. This can be seen from Fig. 4(c) and (d) which show respectively the pDOS for the C and O 2p states of the CO$_2$ molecule and the 4d states of the Nb adatom for the situation where CO$_2$ is far (10 Å) from the surface (before bonding) and in the equilibrium geometry (after bonding). Finally, from Fig. 4(c), it can be seen that the two characteristic $\pi_g$ and $\pi_u$ peaks of CO$_2$ (ref. 38) are shifted to higher energy on adsorption. Adsorption of CO$_2$ also causes the spin-up 4d peak of Nb to move above the Fermi level, and the emergence of a spin-down hybridized Nb-4d–O-2p state (cf. Fig. 4(d)). This indicates that spin-up electrons were transferred from Nb to the CO$_2$ molecule by gaining electron charge, with less charge gained from the O-related 2p states of CO$_2$ hybridized with the 4d states of Nb/SV. The magnetization of the different systems is given in Fig. 1 in terms of the total spin magnetic moment of the system before and after adsorption of CO$_2$. An interesting observation is that CO$_2$ adsorption affects the magnetic moment of some systems, with the largest magnetic moment change happening in Nb/DV (total quenching of magnetic moment), Fe/SV (87% drop), Co/SV (80% drop) and Ca/DV (50% drop). However, the Fe/DV, Mn/SV and Mn/DV systems, which have the highest magnetic moments (3.76 $\mu_B$, 2.95 $\mu_B$ and 3.42 $\mu_B$, respectively), do not undergo any significant change in magnetization. In spite of the high reactivity of Ca/SV towards CO$_2$, CO$_2$ only mildly affects its spin. Strong CO$_2$ capture systems Zr/SV, Zr/DV, Ti/SV and Y/SV are non magnetic, and do not become magnetic upon CO$_2$ adsorption. The changing of magnetic moments upon CO$_2$ adsorption highlights the interaction between the adsorbed CO$_2$ molecules and the dopant–vacancy complexes.

Charge transfer has been a key concept to study the interaction between gas molecules and metallic ions. Charge transfer analyses, the free CO$_2$ molecule accepts $-0.27|e|$, which means that the CO$_2$ molecule hole-dopes the Nb/SV system.

Change transfer has been a key concept to study the interaction between gas molecules and metallic ions. Fig. 5 displays the progressive change of the net charge on the metal dopant relative to the metal without CO$_2$ present, for increasing numbers of adsorbed CO$_2$. All of the adatoms in the figure, except Nb/SV system, react to the adsorption of the first CO$_2$ molecule by gaining electron charge, with less charge gained per CO$_2$ as more CO$_2$ molecules are adsorbed. This behaviour is in contrast to the case of hydrogen adsorption, reported in ref. 18, where Ca, Sc, Ti, Y and Zr adsorbed on DV graphene interacts with H$_2$ molecules by donating charge. Upon adsorption of the first CO$_2$, the metal dopant in Zr/DV acquires the greatest amount of electron charge (0.256e), followed by the metal dopant in Sc/SV (0.204e), Ti/SV (0.202e), Y/SV (0.191e), and Zr/SV (0.184e), while in the rest of the systems the metal adatom acquire a much smaller amount of electron charge. As the systems reach the maximum capacity, the metal adatom in Y/SV has accumulated the largest amount of charge (0.341e). In spite of displaying greater adsorption energies than other systems, Nb/DV and Nb/SV display only a small net charge transfer. In the Voronoi formalism, the first adsorbed CO$_2$ in Ca/SV, Ca/DV, Co/SV, Fe/SV, Nb/SV, Nb/DV and Ni/SV receives electrons, whereas the first CO$_2$ adsorbed on the rest of the systems loses electron charge. All of the aforementioned 7 systems in addition to Zr/SV involve the creation of the CO$_3$ anion (i.e. the bending of the CO$_2$ molecule, see Sec. 3.3). Upon the adsorption of CO$_2$, the metal dopant in Zr/DV acquires the greatest amount of electron charge (0.256e), followed by the metal dopant in Sc/SV (0.204e), Ti/SV (0.202e), Y/SV (0.191e), and Zr/SV (0.184e), while in the rest of the systems the metal adatom acquire a much smaller amount of electron charge.
of subsequent CO$_2$, the nCO$_2$ complex for these systems donates electron charge to the metal–graphene system.

The variation of charge on the dopant ion and CO$_2$ molecules is related to the defective graphene sheet. The charge transfer and redistribution processes among the components of each gas–metal–graphene system exhibits vast diversity, as shown in the schematic diagram in Fig. 6. The diagram shows that the edge C atoms surrounding the defect, as well as the whole graphene sheet, play an important role in the reaction with the adsorbed CO$_2$ molecules. Before adsorbing the first CO$_2$ charge has been transferred from the metal adatom to the graphene sheet, mostly into the C$_e$ atoms. Generally, the charge transfer process involved in the adsorption of the first CO$_2$ is dramatically different from that involved in the adsorption of subsequent CO$_2$ molecules. The CO$_2$ complex tends to keep on losing charge, while the graphene sheet tends to keep on attracting charge as more CO$_2$ molecules are adsorbed. This highlights a novel role of the defective graphene, where it plays the role of a charge reservoir for binding CO$_2$ molecules. Except for Ca/SV, starting from the second adsorbed CO$_2$ molecule, the metal acts as a “channel” for charge transfer, where charge passes through the metal from the CO$_2$ complex to the graphene charge reservoir. In general, SV systems donate more charge to the first adsorbed CO$_2$ than DV systems.

To gain further insight into charge distribution, we plot the density difference (Δρ(r)) distributions for two representative systems, Ca/DV and Sc/SV with five adsorbed CO$_2$ molecules, in Fig. 7, where Δρ is given by

$$\Delta \rho(r) = \rho_{nCO_2}(r)_{M/SV(DV)/Gr} - \rho_{nCO_2}(r)_{M/SV(DV)/Gr} - \rho_{nCO_2}(r)_{Gr}.$$  (5)

Here $\rho_{nCO_2}(r)_{M/SV(DV)/Gr}$ is the charge density of the system M/SV(DV) with nCO$_2$ molecules adsorbed, $\rho_{M/SV(DV)/Gr}$ is the electron charge density of the M/SV(DV) system with no adsorbed CO$_2$ molecules, and $\rho_{nCO_2}(r)_{Gr}$ is the charge density computed for nCO$_2$ molecules at the same coordinates as those in the M/SV(DV) system with adsorbed nCO$_2$ molecules.

In Fig. 7(a), we can observe that the metal adatom gains charge, which is consistent with the series of green triangles for Ca/DV in Fig. 6. The Sc adatom in the Sc/SV system (Fig. 7(b)) is surrounded by smaller blue and red clouds (small charge densities), corresponding to the series of yellow circles for Sc/SV in Fig. 6. The difference between the two systems is that the graphene sheet is more involved in the electron charge redistribution for Ca/DV than for Sc/SV. An interesting observation in both systems is that the interaction between the adatom and the CO$_2$ molecule involves a polarization of the CO$_2$ molecule. This can be seen in Fig. 7 where the oxygen atom of CO$_2$ closest to the adatom loses charge, while the one further away from the adatom gains charge.

### 3.3 Formation of bent CO$_2$ anions

Interestingly, the charge transfer also has an impact on the structure of the first adsorbed CO$_2$ molecule. The D $\neq$ h symmetry of CO$_2$ of the first CO$_2$ molecule adsorbed to seven of the systems considered, namely Ca, Nb, Co, Ni, Fe, Zr and Y, is broken due to significant metal adatom–CO$_2$ molecule interaction. This interaction, for Ca, Nb, Co, Ni and Fe, involves considerable charge transfer to CO$_2$. For these systems, the bending of the CO$_2$ molecule occurs without an energy barrier. This results in weakened and longer C–O bonding, as well as an unbalanced charge occupation on the two O atoms of the CO$_2$ molecule. In all of the systems that experience O–C–O bending, the C–O bond length increases from the initial 1.18 Å. The O–C–O angles and C–O bond lengths in the non-collinear cases are: 131.6°, 1.26 Å in Ca/DV, 130.5°, 1.26 Å in Ca/SV, 152.5°, 1.22 Å in Co/SV, 151.6°, 1.23 Å in Nb/DV, 140.9°, 1.29 Å in Nb/SV, 148.6°, 1.26 Å in Ni/SV, 147°, 1.27 Å in Fe/SV, 155.0°, 1.24 Å in Y/SV and 154.2°, 1.25 Å in Zr/DV. These CO$_2$ anions could potentially react with two protons, to generate a CO and H$_2$O.

For the Ca and Nb systems, the first adsorbed CO$_2$ molecule gains significant electronic charge, 0.29e for Ca/DV, 0.34e for Ca/
SV, 0.23e for Nb/DV and 0.27e for Nb/SV, which are larger than the respective values in other systems. This explains why the first adsorbed CO2 for these systems bend by larger angles than the rest of the systems. These two adatoms bond with the three atoms of the first CO2 molecule. This is reminiscent of the formation of metal carbonates, which are known to form at many metal surfaces. For the Ni, Co, Zr and Y adatoms, they bond with the C atom of CO2 and one of the two O atoms. Note that the bending is greatest in the strongest two adsorbents, Ca and Nb, but the angle also significantly changes in the case of a weak adsorbent, namely Ni.

Metal elements embedded in SV and DV graphene have distinct chemical behavior from their bulk counterparts. The formation of nickel carbonate was reported by Freund and Roberts in their experimental review work. It is interesting to note that some metals that form bulk carbonates in the presence of CO2 without requiring additional energy, oxidation or catalysis fail to do so when adsorbed on SV/DV graphene. For example, CO2 forms a carbonate on the Cu(211) surface, on the Mg(0001) surface and on Al foil.

### 4 Conclusions

We report a comprehensive first-principles DFT study of multiple CO2 capture on 16 representative metal-doped graphene systems, for the dopant adsorbed on single and double vacancies (SV and DV, respectively) in graphene. We use a combination of adsorption energy and bond distance criteria to evaluate the CO2 adsorption capacity. Significantly, our results reveal the dual beneficial roles of vacancies in graphene: (1) to effectively stabilize the adsorbed metal adatom, and (2) to enable defective graphene to act as a charge reservoir for enhancing the CO2 capture capacity. Generally, SV graphene with metal dopants are better sorbents than DV systems. Importantly, our results predict that each Ca, Sc and Y adatom can capture up to 5 CO2 molecules. However, in spite of displaying superior capture capacity, Ca systems readily create CaO upon reaction with CO2, thus rendering Ca/SV and CA/DV graphene less reusable for CO2 capture. Also interestingly, bent CO2 anion molecules are predicted for Ca, Nb, Co, Ni, Fe, Zr and Y doped systems. These findings will be useful in the design and optimization of metal-doped graphene systems for energy and environmental applications.

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