# Nanomaterials Design for Lithium-Sulfur Batteries

A thesis presented for the award of the degree of

# **Doctor of Philosophy**

From

**University of Technology Sydney** 

By

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June, 2020

CERTIFICATE OF ORIGINAL AUTHORSHIP

I, Yi Chen declare that this thesis, is submitted in fulfilment of the requirements for the award

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Science at the University of Technology Sydney.

This thesis is wholly my own work unless otherwise reference or acknowledged. In addition,

I certify that all information sources and literature used are indicated in the thesis.

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This research is supported by the Australian Government Research Training Program.

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Ι

# **DEDICATION**

This thesis is dedicated to my family. Appreciate their love and support.

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- (1) <u>Y. Chen</u>, D. Su, Q. Zhang, G. Wang, 60 Years of Lithium-Sulfur Batteries: From Academic Research to Commercial Viability, *Adv. Mater.*, Submitted.
- (2) <u>Y. Chen</u>, S. Choi, D. Su, X. Gao, G. Wang, Self-standing sulfur cathodes enabled by 3D hierarchically porous titanium monoxide-graphene composite film for high-performance lithium-sulfur batteries, *Nano Energy*, 2018, 47, 331-339. (IF=15.548)
- (3) Y. Chen, W. Zhang, D. Zhou, H. Tian, D. Su, C. Wang, D. Stockdale, F. Kang, B. Li, G. Wang, Co-Fe Mixed Metal Phosphide Nanocubes with Highly Interconnected-Pore Architecture as an Efficient Polysulfide Mediator for Lithium-Sulfur Batteries, ACS Nano, 2019, 13, 4731-4741. (IF=13.903)
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cathode during discharge-charge process. Reproduced from reference 113. Copyright 2019, Wiley-VCH24 <b>Figure 1.8</b> (a) Schematic of the synthesis process for sulfur-TiO <sub>2</sub> yolk-shell nanostructure. (b) SEM and (c  TEM images of yolk-shell structured sulfur-TiO <sub>2</sub> spheres. (d) Cycling performance of sulfur-TiO <sub>2</sub> yolk-shell sphere cathode at 0.5C for 1000 cycles. (a-d) Reproduced from reference 115. Copyright 2013,  Nature Publishing Group. (e) XPS of S/MnO <sub>2</sub> nanosheets electrodes after discharge to certain states: from  top to bottom: discharged to 2.15 V, discharged to 2.15 V and then aged in the cell for 20h, discharged to  800 mA h g <sup>-1</sup> , and discharged to 1.8 V. Reproduced from reference 121. Copyright 2015, Nature Publishing  Group. (f) Schematic of synthesis process of MnO <sub>2</sub> @HCF/S. TEM images of g) MnO <sub>2</sub> @SiO <sub>2</sub> @C, (h)  MnO <sub>2</sub> @HCF and i) MnO <sub>2</sub> @HCF/S at 0.2C. f-k) Reproduced from reference 122. Copyright 2015,  Wiley-VCH	
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TEM images of yolk-shell structured sulfur-TiO <sub>2</sub> spheres. (d) Cycling performance of sulfur-TiO <sub>2</sub> yolk-shell sphere cathode at 0.5C for 1000 cycles. (a-d) Reproduced from reference <sup>115</sup> . Copyright 2013, Nature Publishing Group. (e) XPS of S/MnO <sub>2</sub> nanosheets electrodes after discharge to certain states: from top to bottom: discharged to 2.15 V, discharged to 2.15 V and then aged in the cell for 20h, discharged to 800 mA h g <sup>-1</sup> , and discharged to 1.8 V. Reproduced from reference <sup>121</sup> . Copyright 2015, Nature Publishing Group. (f) Schematic of synthesis process of MnO <sub>2</sub> @HCF/S. TEM images of g) MnO <sub>2</sub> @SiO <sub>2</sub> @C, (h) MnO <sub>2</sub> @HCF and i) MnO <sub>2</sub> @HCF/S (j) Illustration of polysulfide trapping by MnO <sub>2</sub> @HCF. (k) Cycling performances of MnO <sub>2</sub> @HCF/S and HCF/S at 0.2C. f-k) Reproduced from reference <sup>122</sup> . Copyright 2015, Wiley-VCH.  27 Figure 1.9 (a) Schematic of the interaction between Li <sub>2</sub> S <sub>4</sub> and Ti <sub>4</sub> O <sub>7</sub> . (b) Photographs of Li <sub>2</sub> S <sub>4</sub> /THF solution before (top) and after (bottom) contact with nothing (1), graphite (2), VC carbon (3) and Ti <sub>4</sub> O <sub>7</sub> (4). (c) High-resolution XPS S 2p spectra of Li <sub>2</sub> S <sub>4</sub> , Li <sub>2</sub> S <sub>4</sub> /Ti <sub>4</sub> O <sub>7</sub> and Li <sub>2</sub> S <sub>4</sub> /VC, respectively. (d) Distribution of sulfur species upon discharge determined by operando XANES. The solid and dashed lines represent Ti <sub>4</sub> O <sub>7</sub> /S-6 and VC/S-6, respectively. (e) Illustration of surface-mediated reduction of S <sub>8</sub> on Ti <sub>4</sub> O <sub>7</sub> . (a-e) Reproduced from reference <sup>128</sup> . Copyright 2014, Nature Publishing Group. XRD patterns of (f) rutile TiO <sub>2</sub> , (i) Magnéli phase Ti <sub>6</sub> O <sub>11</sub> and (i) Ti <sub>4</sub> O <sub>7</sub> , respectively. The insets show the corresponding fast Fourier transform (FFT) diffraction patterns. (f-n) Reproduced from reference <sup>129</sup> . Copyright 2014, American Chemical Society.  31 Figure 1.10 (a) Schematic of the synthetic process for mesoporous Ti <sub>4</sub> O <sub>7</sub> microspheres. SeM images of (b) mesoporous Ti <sub>2</sub> O <sub>2</sub> microspheres and (c) mesoporous Ti <sub>4</sub> O <sub>7</sub> microspheres. a-c) Reproduced from reference <sup>131</sup> . Copyright 2016, Wiley-VCH. (d) Schematic of the synthetic procedure for Ti <sub>4</sub> O <sub>7</sub> nanop	
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(c) High-resolution XPS S 2p spectra of Li <sub>2</sub> S <sub>4</sub> , Li <sub>2</sub> S <sub>4</sub> /Ti <sub>4</sub> O <sub>7</sub> and Li <sub>2</sub> S <sub>4</sub> /VC, respectively. (d) Distribution of sulfur species upon discharge determined by operando XANES. The solid and dashed lines represent Ti <sub>4</sub> O <sub>7</sub> /S-6 and VC/S-6, respectively. (e) Illustration of surface-mediated reduction of S <sub>8</sub> on Ti <sub>4</sub> O <sub>7</sub> . (a-e) Reproduced from reference <sup>128</sup> . Copyright 2014, Nature Publishing Group. XRD patterns of (f) rutile TiO <sub>2</sub> , (i) Magnéli phase Ti <sub>6</sub> O <sub>11</sub> and (l) Ti <sub>4</sub> O <sub>7</sub> , respectively. The insets are the corresponding stackings of the oxygen octahedral. TEM images of g) TiO <sub>2</sub> , (j) Ti <sub>6</sub> O <sub>11</sub> and (m) Ti <sub>4</sub> O <sub>7</sub> , respectively. The insets are the corresponding digital photos. HR-TEM images of (h) TiO <sub>2</sub> , k) Ti <sub>6</sub> O <sub>11</sub> and n) Ti <sub>4</sub> O <sub>7</sub> , respectively. The insets show the corresponding fast Fourier transform (FFT) diffraction patterns. (f-n) Reproduced from reference <sup>129</sup> . Copyright 2014, American Chemical Society	
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<b>Figure 1.10</b> (a) Schematic of the synthetic process for mesoporous Ti <sub>4</sub> O <sub>7</sub> microspheres. SEM images of (b) mesoporous TiO <sub>2</sub> microspheres and (c) mesoporous Ti <sub>4</sub> O <sub>7</sub> microspheres. a-c) Reproduced from reference <sup>130</sup> . Copyright 2016, Wiley-VCH. (d) Schematic of the synthetic procedure for Ti <sub>4</sub> O <sub>7</sub> nanoparticles with interconnected-pore architecture. TEM images of (e) porous PS-P2VP particles, (f) TiO <sub>2</sub> @PS-P2VP particles and (g) porous Ti <sub>4</sub> O <sub>7</sub> particles. (d-g) Reproduced from reference <sup>131</sup> . Copyright 2017, Wiley-VCH	
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TiO <sub>2</sub> @PS-P2VP particles and (g) porous Ti <sub>4</sub> O <sub>7</sub> particles. (d-g) Reproduced from reference <sup>131</sup> . Copyright 2017, Wiley-VCH	reference <sup>130</sup> . Copyright 2016, Wiley-VCH. (d) Schematic of the synthetic procedure for Ti <sub>4</sub> O <sub>7</sub>
2017, Wiley-VCH	nanoparticles with interconnected-pore architecture. TEM images of (e) porous PS-P2VP particles, (f)
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TiO@C-HS/S cathodes at 0.2 and 0.5C. (a-f) Reproduced from reference <sup>134</sup> . Copyright 2016, Nature Publishing Group. (g) Schematic of synthesis procedure of 3D hierarchically porous TiO-G/S film. Digital photos of (b) PS-TiO <sub>2</sub> -GO sponge and (i) TiO-G film electrodes. (j) Cross-sectional backscattered electron image of sulfur-loaded TiO-graphene film. k) Illustration of the interactions between TiO and sulfur species. (g-k) Reproduced from reference <sup>135</sup> . Copyright 2018, Elsevier	Figure 1.11 (a) Schematic of synthetic process for TiO@C-HS/S spheres. TEM images of (b) PS, (c)
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species. (g-k) Reproduced from reference <sup>135</sup> . Copyright 2018, Elsevier	• • • • • • • • • • • • • • • • • • • •
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#### **ABSTRACT**

Lithium-sulfur (Li-S) batteries, which rely on the redox reactions, show great promise for next-generation energy storage owing to their high theoretical energy density, environmental benignity and low cost of sulfur. However, the practical application of Li-S batteries has been largely impeded by the low conductivity of sulfur and the shuttle effect of polysulfides. One of the most effective strategies to overcome these problems is to disperse insulating sulfur active material within other conductive matrixes that are capable of physically adsorbing and/or chemically binding sulfur and its intermediate polysulfides. In this thesis, we designed two types of host materials that can be used to improve the electrochemical performance of Li-S batteries.

A new self-standing host enabled by a 3D hierarchically-porous titanium monoxide-graphene composite film was designed to overcome the main challenges of Li-S batteries. The hierarchically porous graphene scaffold can not only facilitate rapid lithium ion and electron transport, but also provide sufficient spaces to accommodate sulfur species. In addition, the ultrafine and polar titanium monoxide nanoparticles embedded in the three-dimensional graphene networks show strong chemical anchoring for polysulfides, and their inherent metallic conductivity accelerates the redox reaction kinetics. Benefiting from this attractive architecture, the freestanding titanium monoxide-graphene/sulfur cathode demonstrated superior electrochemical performance for Li-S batteries.

Uniform Co-Fe mixed metal phosphide (Co-Fe-P) nanocubes with highly interconnected-pore architecture were synthesized as sulfur host for Li-S batteries. With the

highly interconnected-pore architecture, inherently metallic conductivity and polar characteristic, the Co-Fe-P nanocubes not only offer sufficient electrical contact to the insulating sulfur for high sulfur utilization and fast redox reaction kinetics, but also provide abundant adsorption sites for trapping and catalyzing the conversion of lithium polysulfides to suppress the shuttle effect. As a result, the sulfur-loaded Co-Fe-P (S@Co-Fe-P) nanocubes exhibited superior electrochemical performances both in coin cells and pouch cells.